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Chapter 11 - Stoichiometry Section 11.1 What
is stoichiometry? 1. true 2. true 3. false 4.
true 5. true 6. 2, 2, 64.10 7. 3, 3, 96.00 8.
2, 2, 88.02 9. 4, 4, 72.08 10. methanol and
oxygen gas 11. carbon dioxide and water 12.
160.10 g 13. 160.10 g 14. They are equal. 15.
A mole ratio is a ratio between the numbers
of moles

VIBRATIONS AND WAVES - Simon Technology
368 Chapter 11 • Stoichiometry Section
11.1.1 Objectives Describe the types of
relationships indicated by a balanced
chemical equation. State the mole ratios from
a balanced chemical equation. Review
Vocabulary reactant: the starting substance
in a chemical reaction New Vocabulary
stoichiometry mole ratio Defining
Stoichiometry

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Stoichiometry. stoichiometry. mole ratio.
limiting reactant. excess reactant. the study
of quantitative relationships between the
amounts of... in a balanced equation, the
ratio between the number of moles... a
reactant that is totally consumed during a
chemical reaction...

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CHAPTER 11 SOLUTIONS MANUAL Section 11.1

Defining Stoichiometry pages 368–372 Practice

Problems pages 371–372 1. Interpret the

following balanced chemical equa-tions in

terms of particles, moles, and mass. Show

that the law of conservation of mass is

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Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the amounts of products formed by a chemical reaction

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In Section 11.3 , for example, you learned how to express the stoichiometry of the reaction for the ammonium dichromate volcano in terms of the atoms, ions, or molecules involved and the numbers of moles, grams, and formula units of each (recognizing, for instance, that 1 mol of ammonium dichromate produces 4 mol of water). This section describes how to use the stoichiometry of a reaction to answer questions like the following: How much oxygen is needed to ensure complete combustion of a ...

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