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Solution: $p_{abs} = p_{atm} + p_{gage} = 101:3kPa + 20 \text{ in. H}_2\text{O}$
 $248:8Pa$ 1:0 in. $\text{H}_2\text{O} = 110:26kPa$ Ideal gas law: $\rho = \frac{p}{RT} =$
 $110:26kPa / (1 \text{ kgK} \cdot 2077J / 1 \text{ 293:2K} \cdot 1000Pa / 1kPa) = 1 \text{ Nm}^{-3}$
 $\rho = 0:181 \text{ kg/m}^3$ b.) Situation: A sphere of 93 mm diameter contains an
ideal gas. $T = 20 \text{ C} = 293:2K$ Find: Calculate the density of argon at
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